

**GCE** 

# **Chemistry A**

H432/02: Synthesis and analytical techniques

A Level

Mark Scheme for June 2022

OCR (Oxford Cambridge and RSA) is a leading UK awarding body, providing a wide range of qualifications to meet the needs of candidates of all ages and abilities. OCR qualifications include AS/A Levels, Diplomas, GCSEs, Cambridge Nationals, Cambridge Technicals, Functional Skills, Key Skills, Entry Level qualifications, NVQs and vocational qualifications in areas such as IT, business, languages, teaching/training, administration and secretarial skills.

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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#### MARKING INSTRUCTIONS

#### PREPARATION FOR MARKING

#### RM ASSESSOR

- 1. Make sure that you have accessed and completed the relevant training packages for on-screen marking: *RM Assessor Online Training*; *OCR Essential Guide to Marking*.
- 2. Make sure that you have read and understood the mark scheme and the question paper for this unit.
- 3. Log-in to RM Assessor and mark the **required number** of practice responses ("scripts") and the **required number** of standardisation responses.

#### **MARKING**

- 1. Mark strictly to the mark scheme.
- 2. Marks awarded must relate directly to the marking criteria.
- 3. The schedule of dates is very important. It is essential that you meet the RM Assessor 50% and 100% (traditional 50% Batch 1 and 100% Batch 2) deadlines. If you experience problems, you must contact your Team Leader (Supervisor) without delay.
- 4. If you are in any doubt about applying the mark scheme, consult your Team Leader by telephone, email or via the RM Assessor messaging system.
- Work crossed out:

#### **Crossed Out Responses**

Where a candidate has crossed out a response and provided a clear alternative then the crossed out response is not marked. Where no alternative response has been provided, examiners may give candidates the benefit of the doubt and mark the crossed out response where legible.

#### **Rubric Error Responses – Optional Questions**

Where candidates have a choice of question across a whole paper or a whole section and have provided more answers than required, then all responses are marked and the highest mark allowable within the rubric is given. Enter a mark for each question answered into RM assessor, which will select the highest mark from those awarded. (The underlying assumption is that the candidate has penalised themselves by attempting more questions than necessary in the time allowed.)

## **Multiple Choice Question Responses**

When a multiple choice question has only a single, correct response and a candidate provides two responses (even if one of these responses is correct), then no mark should be awarded (as it is not possible to determine which was the first response selected by the candidate). When a question requires candidates to select more than one option/multiple options, then local marking arrangements need to ensure consistency of approach.

## **Contradictory Responses**

When a candidate provides contradictory responses, then no mark should be awarded, even if one of the answers is correct.

#### **Short Answer Questions** (requiring only a list by way of a response, usually worth only **one mark per response**)

Where candidates are required to provide a set number of short answer responses then only the set number of responses should be marked. The response space should be marked from left to right on each line and then line by line until the required number of responses have been considered. The remaining responses should not then be marked. Examiners will have to apply judgement as to whether a 'second response' on a line is a development of the 'first response', rather than a separate, discrete response. (The underlying assumption is that the candidate is attempting to hedge their bets and therefore getting undue benefit rather than engaging with the question and giving the most relevant/correct responses.)

## Short Answer Questions (requiring a more developed response, worth two or more marks)

If the candidates are required to provide a description of, say, three items or factors and four items or factors are provided, then mark on a similar basis – that is downwards (as it is unlikely in this situation that a candidate will provide more than one response in each section of the response space.)

#### **Longer Answer Questions** (requiring a developed response)

Where candidates have provided two (or more) responses to a medium or high tariff question which only required a single (developed) response and not crossed out the first response, then only the first response should be marked. Examiners will need to apply professional judgement as to whether the second (or a subsequent) response is a 'new start' or simply a poorly expressed continuation of the first response.

6. Always check the pages (and additional objects if present) at the end of the response in case any answers have been continued there. If the candidate has continued an answer there then add a tick to confirm that the work has been seen.

- 7. Award No Response (NR) if:
  - there is nothing written in the answer space.

Award Zero '0' if:

• anything is written in the answer space and is not worthy of credit (this includes text and symbols).

Team Leaders must confirm the correct use of the NR button with their markers before live marking commences and should check this when reviewing scripts.

- 8. The RM Assessor **comments box** is used by your Team Leader to explain the marking of the practice responses. Please refer to these comments when checking your practice responses. **Do not use the comments box for any other reason.** 
  - If you have any questions or comments for your Team Leader, use the phone, the RM Assessor messaging system, or email.
- 9. Assistant Examiners will send a brief report on the performance of candidates to their Team Leader (Supervisor) via email by the end of the marking period. The report should contain notes on particular strengths displayed as well as common errors or weaknesses. Constructive criticism of the question paper/mark scheme is also appreciated.

10. For answers marked by levels of response:

Read through the whole answer from start to finish, using the Level descriptors to help you decide whether it is a strong or weak answer. The indicative scientific content in the Guidance column indicates the expected parameters for candidates' answers, but be prepared to recognise and credit unexpected approaches where they show relevance. Using a 'best-fit' approach based on the skills and science content evidenced within the answer, first decide which set of level descriptors, Level 1, Level 2 or Level 3, best describes the overall quality of the answer.

Once the level is located, award the higher or lower mark:

The higher mark should be awarded where the level descriptor has been evidenced and all aspects of the communication statement (in italics) have been met.

**The lower mark** should be awarded where the level descriptor has been evidenced but aspects of the communication statement (in italics) are missing.

## In summary:

The skills and science content determines the level.

The communication statement determines the mark within a level.

Level of response questions on this paper are 19(e) and 21

The only annotation on a level of response question should be the indication of the level.

A level annotation should be used where all marks for a level have been achieved. e.g. if a candidate has 6 marks, they would have this annotation on their script:

L3

If a candidate has achieved 5 marks then they have reached Level 3 but will not have met the communication statement. They should have the following annotations on their scripts:



The same principle should be applied to Level 2 and Level 1.

No marks (0) should have a cross:

Place the annotations alongside the mark for the question.

On additional pages, annotate using

## 11. Annotations available in RM Assessor

Annotation	Meaning
<b>✓</b>	Correct response
×	Incorrect response
^	Omission mark
BOD	Benefit of doubt given
CON	Contradiction
RE	Rounding error
SF	Error in number of significant figures
ECF	Error carried forward
L1	Level 1
L2	Level 2
L3	Level 3
NBOD	Benefit of doubt not given
SEEN	Noted but no credit given
I	Ignore
BP	Blank page

12. Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
_	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

## 13. Subject-specific Marking Instructions

#### INTRODUCTION

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

# **SECTION A**

Question	Answer	Marks	AO element	Guidance
1	В	1	AO1.1	
2	В	1	AO1.1	
3	D	1	AO1.2	ALLOW 15 (correct number of sigma bonds)
4	Α	1	AO1.1	
5	D	1	AO1.2	ALLOW 8 (correct number of chiral centres)
6	В	1	AO1.2	
7	D	1	AO1.2	
8	С	1	AO2.2	ALLOW 500 (This is the correct mass)
9	С	1	AO2.6	ALLOW 4.8 (This is the correct volume)
10	В	1	AO1.2	
11	В	1	AO2.5	
12	Α	1	AO2.1	
13	D	1	AO1.1	
14	Α	1	AO1.1	
15	В	1	AO2.1	
	Total	15		

# **SECTION B**

(	Question		Answer	Marks	AO element	Guidance
16	(a)	(i)	3-methylhex-2-ene ✓	1	AO1.2	IGNORE lack of hyphens, or addition of commas  DO NOT ALLOW 3-methyhex-2-ene OR 3-methhex-2-ene OR 3-methlyhex-2-ene OR 3-methylhexan-2-ene IGNORE references to E/Z or cis/trans
16	(a)	(ii)	ANNOTATE ANSWER WITH TICKS AND CROSSES $H_3C \qquad CH_3$ $CH_3CH_2CH_2 \qquad H$ $Br\delta^+$ $Curly arrow from C=C bond to Br^{\delta^+} of Br–Br  AND  Correct dipole on Br–Br AND curly arrow for breaking of Br–Br bond \checkmark$	3	AO1.2 ×1 AO2.5 ×2	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  IGNORE connectivity of CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> and CH <sub>3</sub> groups in carbocation and product ALLOW C <sub>3</sub> H <sub>7</sub> for CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> DO NOT ALLOW half headed or double headed arrows but allow ECF if seen more than once  DO NOT ALLOW use of HBr but ECF for subsequent use  For curly arrows, ALLOW straight or snake-like arrows and small gaps (see examples):

Question	Answer		AO element	Guidance	
	Correct carbocation to match mechanism  AND curly arrow from Br to C+ of carbocation  H <sub>3</sub> C  CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> Br  OR  i.e. ALLOW carbonium + on either C atom			2nd curly arrow must  • start from, OR be traced back to, any part  of *Br-Br* bond  • AND go to Br*  Br*  Br*  Br*  Br*  AND go to Br*  Gr*  Gr*  Gr*  Gr*  Gr*  Gr*  Gr*	

C	Question		Answer		AO element	Guidance	
			CH <sub>3</sub> CH <sub>3</sub> CH <sub>3</sub> CH <sub>3</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> Br Br				
16	(b)	(i)	Same molecular formula  AND  Different structural formulae ✓  OR  Both have the molecular formula C <sub>6</sub> H <sub>12</sub> AND  Different structural formulae ✓	1	AO1.1	Same formula is <b>not</b> sufficient	
16	(b)	(ii)	Same structural formula  AND  Different arrangement (of atoms) in space  OR different spatial arrangement (of atoms) ✓	1	AO1.1	ALLOW structure/displayed/skeletal formula  DO NOT ALLOW same empirical formula OR same general formula  IGNORE same molecular formula  Reference to E/Z isomerism or optical isomerism is not sufficient	

Question	Answer	Marks	AO element	Guidance
Question  16 (b) (iii)	Correct identification of <i>cis</i> <b>AND</b> <i>trans</i> isomers of  4-methylpent-2-ene   H  H  CH(CH <sub>3</sub> ) <sub>2</sub> Cis isomer  Trans isomer  OR  Identification of 3-methylpent-2-ene as <i>cis</i> <b>AND</b> <i>trans</i> isomers  V  CH <sub>2</sub> CH <sub>3</sub> H  CH <sub>2</sub> CH <sub>3</sub> H  CH <sub>2</sub> CH <sub>3</sub> CH <sub>3</sub> C  CH <sub>3</sub>	Marks 2		Guidance  ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  C <sub>3</sub> H <sub>7</sub> is <b>not</b> sufficient (could be unbranched)  ALLOW one mark if <i>cis</i> AND <i>trans</i> isomers of 4-methylpent-2-ene are in the wrong boxes  ALLOW the isomers of 3-methylpent-2-ene in either box  H <sub>3</sub> C  CH <sub>2</sub> CH <sub>3</sub> H  CH <sub>2</sub> CH <sub>3</sub> ALLOW one mark for correct identification system  ALLOW one mark for correct identification of <i>cis</i> AND <i>trans</i> isomers of unbranched C <sub>6</sub> H <sub>12</sub> e.g.  H  CH <sub>3</sub> CH <sub>2</sub> H  CH <sub>3</sub> CH <sub>2</sub> H  CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>
				cis isomer trans isomer

(	Quest	ion	Answer	Marks	AO element	Guidance
16		(iv)	Correct groups attached to chiral carbon of compound C seen <b>once</b> e.g.  CH=CH <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	Marks 2		ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  For C <sub>2</sub> H <sub>5</sub> -, ALLOW CH <sub>3</sub> CH <sub>2</sub> - For -CH=CH <sub>2</sub> , ALLOW -C <sub>2</sub> H <sub>3</sub> OR -CHCH <sub>2</sub> For bond into paper accept:   ALLOW two 3D structures with 2 groups swapped e.g.  CH=CH <sub>2</sub> C <sub>2</sub> H <sub>5</sub> CH=CH <sub>2</sub>

	Question		Answer	Marks	AO element	Guidance
16	(b)	(v)	H <sub>3</sub> C CH <sub>3</sub>	4	AO2.5 ×2	ALLOW 1 mark for structures if shown in wrong boxes.
			<ul> <li>Two of the following for D ✓</li> <li>All H are equivalent/in the same chemical environment/ the same type</li> <li>All C are equivalent/ in the same chemical environment/ the same type</li> <li>No C=C present</li> </ul>		AO2.2 ×2	CHECK table 16.1 for annotations that may be worthy of credit
			<ul> <li>Two of the following for E ✓</li> <li>All H are equivalent/ in the same chemical environment/ the same type</li> <li>2 C environments</li> <li>C=C present</li> </ul>			

(	Quest	ion	Answer	Marks	AO element	Guidance
16	(c)	(i)	BOTH structures required for ✓	2	AO3.1 ×1	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous
					AO3.2 ×1	
16	(c)	(ii)	$H_3C$ $C$ $C$ $C$ $C$ $C$ $C$ $C$ $C$ $C$	1	AO3.2	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous
			Total	17		

Question	Answer	Marks	AO element	Guidance
17	FIRST CHECK ANSWER LINES If M=168(.0) Award 4 marks for calculation providing unit conversions are correct  Use of ideal gas equation $pV = nRT \ OR \ n = \frac{pV}{RT} \ \checkmark$ SI Unit conversions AND substitution into $n = \frac{pV}{RT}$ :  • $R = 8.314 \ OR \ 8.31$ • $V = 186 \times 10^{-6}$ • $T \ in \ K$ : $303 \ K$ e.g. $\frac{1.07 \times 10^5 \times 186 \times 10^{-6}}{8.314 \times 303}$ Calculation of $n$ $n = 7.90 \times 10^{-3} \ (\text{mol}) \ \checkmark$ Calculation of $M$ $M = \frac{1.327}{7.90 \times 10^{-3}} = 168(.0) \ \checkmark$	6	AO1.2 ×1 AO2.4 ×3 AO2.5 ×2	ALLOW ECF throughout  ALLOW calculator value of 167.968115 (using 8.314) for M  ALLOW calculator value of 167.8873033 (using 8.31) for M  Calculator value of <i>n</i> :  from 8.314 = 7.900308915 × 10 <sup>-3</sup> from 8.31 = 7.904111711 × 10 <sup>-3</sup>
	Molecular formula C <sub>3</sub> H <sub>2</sub> F <sub>6</sub> O ✓			ALLOW ECF that matches M but the formula MUST contain $F_6O$ Use of 24 dm <sup>3</sup> : e.g. $n = \frac{186.0}{24000} = 7.75 \times 10^{-3}$ No mark (calculation much simpler) $M = \frac{1.327}{7.75 \times 10^{-3}} = 171(.2) \checkmark \text{ ECF}$ $C_3H_5F_6O \checkmark \text{ ECF}$

Question	Answer	Marks	AO element	Guidance
	Structure F F F OR			ALLOW ECF for a feasibile chemical structure that matches M AND contains F <sub>6</sub> O AND has a chiral carbon  DO NOT ALLOW  F F H H
	F F F			F F no chiral carbon
	Total	6		

C	uesti	ion	Answer	Marks	AO element	Guidance
18	(a)	(i)	(Add) 2,4-dinitrophenylhydrazine <b>AND</b> orange/yellow/red precipitate ✓	3	AO1.2 × 3	ALLOW errors in spelling ALLOW 2,4(-)DNP OR 2,4(-)DNPH ALLOW Brady's reagent or Brady's Test ALLOW solid OR crystals OR ppt as alternatives for precipitate
			Take melting point (of crystals) ✓ Compare to known values/database ✓			Mark second and third points independently of response for first marking point  DO NOT ALLOW 2 <sup>nd</sup> and 3 <sup>rd</sup> marks for taking and comparing boiling points OR chromatograms
18	(a)	(ii)	Tollens' (reagent) AND Silver (mirror/precipitate/ppt/solid) ✓	1	AO1.2	ALLOW ammoniacal silver nitrate OR Ag <sup>+</sup> /NH <sub>3</sub> ALLOW black ppt OR grey ppt  ALLOW Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup> /H <sup>+</sup> AND Turns green ✓  IGNORE reference to conditions, e.g. Heat or reflux

Question	Answer	Marks	AO element	Guidance
18 (b)	Initial ratios $C, \frac{62.07}{12.0}: H, \frac{10.34}{1.0}: O, \frac{27.59}{16.0}$ OR $C, 5.1725: H, 10.34: O, 1.724$	4		CHECK spectra for annotations that may be worthy of credit
	OR C, 3 : H, 6 : O, 1 ✓		AO1.2	Mark can be awarded from a correct molecular formula
	(Molecular formula =) C <sub>3</sub> H <sub>6</sub> O		AO2.5	
	Evidence of <b>58</b> in working or from labelled peak in one of the spectra ✓		AO3.2	
	For F evidence for fragment ion m/z=29 linked to CH₃CH₂(⁺) OR CHO(⁺) ✓		AO3.2	IGNORE m/z=15 (as this is not unique) IGNORE m/z=43  IGNORE incorrect fragments IGNORE charges on fragment ions
	<b>F</b> =O			ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous
	AND  G =			

Q	uesti	on	Answer	Marks	AO element	Guidance
18	(c)		<sup>1</sup> H NMR  δ = 1.1 ppm/doublet linked to 2 x CH <sub>3</sub> ✓	4	AO3.1 ×3	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  CHECK spectra for annotations that may be worthy of credit  ALLOW δ values ± 0.2 ppm, as a range or a value within the range
			$\delta$ = 2.2 ppm/singlet linked to CH <sub>3</sub> –C=O OR $\delta$ = 2.9 ppm/multiplet linked to CH(CH <sub>3</sub> ) <sub>2</sub> OR HC–C=O $\checkmark$ Structure  Any structure with molecular formula C <sub>6</sub> H <sub>10</sub> O <sub>2</sub> and has 2 carbonyl groups $\checkmark$		AO3.2 ×1	IGNORE HC–C=O linked to $\delta$ = 2.2 ppm  IGNORE additional chemical environments (taken from the data sheet) that align with the given chemical shifts
			¹ O ✓ Total	12		

C	Question	Answer	Marks	AO element	Guidance
19	(a)	$C_2H_5COOH + KOH → C_2H_5COOK + H_2O$ ✓ $2HCOOH + Mg → (HCOO)_2Mg + H_2 ✓$	4	AO2.6 ×4	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  IGNORE state symbols and use of equilibrium sign  ALLOW KC <sub>2</sub> H <sub>5</sub> COO  DO NOT ALLOW a missing charge (e.g. C <sub>2</sub> H <sub>5</sub> COO <sup>-</sup> K) the 1st time seen but IGNORE for next equations.  For salts,  ALLOW C <sub>2</sub> H <sub>5</sub> COO <sup>-</sup> K <sup>+</sup> OR C <sub>2</sub> H <sub>5</sub> COO <sup>-</sup> + K <sup>+</sup> DO NOT ALLOW –COO–K (covalent bond) the 1st time seen but IGNORE for next equations.
		H <sub>2</sub> O <b>AND</b> CO <sub>2</sub> ✓			FOR CO <sub>2</sub> + H <sub>2</sub> O ALLOW H <sub>2</sub> CO <sub>3</sub>
		H H <sub>2</sub> N — C — COONa   CH <sub>2</sub>   COONa ✓			

Question	Answer		AO element	Guidance
19 (b)	$\begin{array}{c} H \\ CH_3CH_2 \\ \hline \\ C \\ C$	4	AO2.5 *4	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  ALLOW any vertical bond to the OH group e.g. ALLOW  OR OH HO  IGNORE connectivity of CH <sub>3</sub> CH <sub>2</sub> group  IGNORE inorganic by-products  ALLOW HCI/H <sub>2</sub> O, H <sub>2</sub> SO <sub>4</sub> /H <sub>2</sub> O IGNORE dilute

C	uestic	on	Answer	Marks	AO element	Guidance
19	(c)	(i)	Pent-3-enoic acid	2	AO1.2	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous ALLOW either the E or Z isomer
			2 repeat units of polymer  CH <sub>2</sub> COOH CH <sub>2</sub> COOH  H  CC  CC  CC  CH <sub>3</sub> H  CH <sub>3</sub> H  CH <sub>3</sub> H		AO2.5	ALLOW ECF from pent-2-enoic acid OR pent-4-enoic acid ONLY  For repeat unit,  • 'side bonds' required on either side of repeat unit from C atoms  • 2 repeat units required  IGNORE connectivity of CH <sub>2</sub> COOH in polymer IGNORE brackets IGNORE n

Q	uestic	on	Answer	Marks	AO element	Guidance
19	(c)	(ii)	OHHHO  ester link ✓  ONE repeat units of correct polymer ✓  HO  OH	1	AO1.2 AO2.5	end –O– may be at either side e.g.  ALLOW CH <sub>3</sub> to be on position 2 or 3 of the aromatic ring  'End bonds' MUST be shown (do not have to be dotted) IGNORE brackets IGNORE n  ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous
19	(d)	(i)	но — он — — — — — — — — — — — — — — — —	1	AO2.5	end –N– may be at either side e.g.  H O
			H ĊH₃ ✓  ONE repeat unit ONLY			CH <sub>3</sub> H 'End bonds' <b>MUST</b> be shown (do not have to be dotted) <b>IGNORE</b> brackets <b>IGNORE</b> n

C	uestic	on	Answer	Marks	AO element	Guidance
19	(d)	(ii)	IF answer on answer line = 28418, AWARD 2 marks IF answer on answer line = 28400, AWARD 1 mark $$	2	AO2.2 ×2	<b>ALLOW ECF</b> from incorrect repeat unit in 19di <b>ALLOW ECF</b> from incorrect $M_r$ of 400 repeat units  Alternative method based on repeat unit: $M_r$ of 400 repeat units = $400 \times 71 = 28400 \checkmark$
19	(e)*		Refer to marking instructions on page 5 of mark scheme for guidance on marking this question.  Level 3 (5-6 marks) Correct calculation of mass of CH <sub>3</sub> CHCICOOH.  AND Planned synthesis includes substitution of –Cl and formation of compound I (or its corresponding ammonium salt) with the correct reagents and some conditions identified and equations are mostly correct.  There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.	6	AO3.3 ×6	$M_{\rm f}$ of polymer = 28400 + 1 + 17 = 28418 ✓  Indicative scientific points may include:  Calculation of mass of CH <sub>3</sub> CHCICOOCH <sub>3</sub> Using moles  • $n(\mathbf{I}) = \frac{9.36}{117.0}$ = 0.08(00) (mol)  • $n(\text{CH}_3\text{CHCICOOC}_2\text{H}_5) = 0.0800 \times \frac{100}{64}$ = 0.125 (mol)  • Mass of CH <sub>3</sub> CHCICOOH = 108.5 × 0.125 = 13.5625 g

Question	Answer	Marks	AO element	Guidance
	Level 2 (3-4 marks) Calculation of mass of CH <sub>3</sub> CHCICOOH is correct AND Planned synthesis includes one step of the synthesis with the correct reagent and some conditions identified and equation is mostly correct OR Calculation of mass of CH <sub>3</sub> CHCICOOH is partly correct AND Planned synthesis includes substitution of –Cl and formation of compound I (or its corresponding ammonium salt) with the correct reagents OR Attempts to calculate mass of CH <sub>3</sub> CHCICOOC <sub>2</sub> H <sub>5</sub> but makes little progress AND Planned synthesis includes substitution of –Cl and formation of compound I (or its corresponding ammonium salt) with the correct reagents and some conditions identified and equations are mostly correct  There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.			<ul> <li>Using mass</li> <li>Theoretical mass of I = 9.36 × 100/64</li> <li>= 14.625 (g)</li> <li>Theoretical n(CH<sub>3</sub>CHCICOOH) = 14.625/117.0</li> <li>= 0.125 (mol)</li> <li>Mass of CH<sub>3</sub>CHCICOOH = 108.5 × 0.125</li> <li>= 13.5625 g</li> <li>ALLOW slip/rounding errors such as errors in M<sub>r</sub>, e.g. use of 107.5 instead of 108.5 for CH<sub>3</sub>CHCICOOH → 13.4375</li> <li>Examples of partly correct calculations</li> <li>Mass = 5.5552 g from 0.0800 × 64/100 × 108.5</li> <li>(% yield inverted)</li> <li>Mass = 8.68 g from 0.0800 × 108.5</li> <li>(% yield omitted)</li> <li>Synthesis: Either order for 2 stages</li> <li>Substitution of -CI → amine:</li> <li>Reagents: (excess) NH<sub>3</sub></li> <li>Condition: ethanol</li> <li>Equation: CH<sub>3</sub>CHCICOOH + 2NH<sub>3</sub> → CH<sub>3</sub>CHNH<sub>2</sub>COOH + NH<sub>4</sub>CIOR</li> <li>CH<sub>3</sub>CHCICOOH + NH<sub>3</sub> → CH<sub>3</sub>CHCICOOH + NH<sub>4</sub>CIOOR</li> <li>CH<sub>3</sub>CHCICOOH + NH<sub>3</sub> → CH<sub>3</sub>CHCICOOH + HCI</li> </ul>

Question	Answer	Marks	AO element	Guidance
	Level 1 (1-2 marks) Calculation of mass of CH₃CHCICOOH is partly correct OR Planned synthesis includes both steps with some of the reagents and conditions identified OR Attempts equations for both steps but these may contain errors OR Describes one step of the synthesis with reagents, conditions and equation mostly correct  There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.  O marks No response or no response worthy of credit.			<ul> <li>Reagents: CH<sub>3</sub>CH<sub>2</sub>OH</li> <li>Conditions: acid (catalyst), e.g. H<sub>2</sub>SO<sub>4</sub> (reflux/heat)</li> <li>Equation: CH<sub>3</sub>CHNH<sub>2</sub>COOH + CH<sub>3</sub>CH<sub>2</sub>OH → CH<sub>3</sub>CHNH<sub>2</sub>COOCH<sub>2</sub>CH<sub>3</sub> + H<sub>2</sub>O</li> <li>OR</li></ul>
	Total	22		

Q	uesti	on	Answer	Marks	AO element	Guidance
20	(a)	(i)	Indicator <b>AND</b> observation of acidity <b>AND</b> No reaction with carbonate ✓	1	AO1.2 ×1	ALLOW  (Add) bromine AND white precipitate ✓  ALLOW  (Add) FeCl₃ AND violet/purple colour ✓
20	(a)	(ii)	Compound J has 6 peaks/environments/types of carbon ✓  Compound K has 5 peaks/environments/types of carbon ✓  Compound L has 8 peaks/environments/types of carbon ✓	3	AO3.2 ×3	IGNORE any numbers shown on structures  IGNORE chemical shifts
20	(a)	(iii)	ANNOTATE ANSWER WITH TICKS AND CROSSES  Action of catalyst 1 mark  Formation of electrophile: $Cl_2 + AlCl_3 \rightarrow Cl^+ + AlCl_4^-$ AND  Regeneration of catalyst: $H^+ + AlCl_4^- \rightarrow AlCl_3 + HCl \checkmark$ Electrophilic attack 1 mark  Curly arrow from $\pi$ -bond to $Cl^+ \checkmark$ $Cl^+ \leftarrow CH_2CH_2OH$	4	AO1.2 ×2 AO2.5 ×2	ALLOW use of FeCl <sub>3</sub> or other halogen carriers (AlBr <sub>3</sub> )  For curly arrows, ALLOW straight or snake-like arrows and small gaps (see examples):  1st curly arrow must  • start from, OR close to circle of benzene ring  AND  • go to Cl <sup>+</sup>

Answer		Marks	AO element	Guidance
Correct intermediate only	1 mark			DO NOT ALLOW the following intermediate:
H_	H <sub>2</sub> OH ✓			π-ring must cover more than half of benzene ring AND correct orientation, <i>i.e.</i> gap towards C with Cl ALLOW + sign anywhere inside the 'hexagon' of intermediate  DO NOT ALLOW intermediates substituted at
				positions 3 or 5 IGNORE intermediates substituted at position 2 OR di-substituted at positions 2,4
Reforming benzene ring	1 mark			Curly arrow must start from, OR be traced back to, any part of C–H bond and go inside the
Н				'hexagon'
	Correct intermediate only  H  CH <sub>2</sub> CH  CH  CH  CH  CH  CH  CH  CH  CH  CH	Correct intermediate only  H  CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> OH  Reforming benzene ring  1 mark  Curly arrow from C–H bond to reform π-ring ✓  H  CH <sub>2</sub> CH <sub>2</sub> OH	Correct intermediate only  H  CH₂CH₂OH  CH₂CH₂OH  Reforming benzene ring  1 mark  Curly arrow from C−H bond to reform π-ring ✓  H  CH₂CH₂OH	Correct intermediate only  1 mark  H CH <sub>2</sub> CH <sub>2</sub> OH  Reforming benzene ring 1 mark  Curly arrow from C–H bond to reform π-ring ✓  H CH <sub>2</sub> CH <sub>2</sub> OH

C	Question		Answer	Marks	AO element	Guidance
20	(b)		(In phenols) a (lone) pair of electrons on O is (partially) delocalised/donated into the ring / π-system ✓  Electron density increases/is higher (than benzene) ✓  ORA  (phenols) are more susceptible to electrophilic attack OR	3	AO1.1 × 3	ALLOW the electron pair in the p-orbitals of the O atom becomes part of the ring / π-system ALLOW diagram to show movement of lone pair into ring ALLOW lone pair of electrons on O is (partially) drawn/attracted/pulled/ into ring / π-system ALLOW lone pair on O DO NOT ALLOW (two) lone pairs are delocalised/donated into the ring / π-system IGNORE activating  IGNORE charge density IGNORE phenols react more readily with
			(phenols) attract/accept electrophile/Cl₂ more  OR  (phenols) polarise electrophile/Cl₂ more ✓  ORA			electrophiles/Cl <sub>2</sub> (given in question)  ALLOW Cl <sup>+</sup> for electrophile  IGNORE Cl for electrophile
20	(c)		Correct balanced equation ✓	2	AO2.5 AO2.6	ALLOW any combination of skeletal <b>OR</b> structural <b>OR</b> displayed formula as long as unambiguous  ALLOW C <sub>6</sub> H <sub>5</sub> for phenyl group
			Correct balanced equation ✓			

20	(d)	(i)	HNO <sub>3</sub> /nitric acid  AND H₂SO <sub>4</sub>	2	AO1.2 ×2	IGNORE references to concentration
			OH Sn AND HCI			IGNORE 'dilute' for HC <i>l</i> IGNORE H₂ IGNORE NaOH if seen as a reagent to convert nitro group into amine e.g 'Sn/(concentrated) HC <i>l</i> then NaOH' scores the mark
20	(d)	(ii)	OH + 6 [H] + 2 H <sub>2</sub> O NH <sub>2</sub>	1	AO2.6	
20	(e)		Stage 1  Reagents: H₂SO₄ ✓	4	AO3.1	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous
			OH + H <sub>2</sub> O		AO2.6	ALLOW H <sup>+</sup> OR HCl OR H <sub>3</sub> PO <sub>4</sub> DO NOT ALLOW other named acids IGNORE concentration/pressure IGNORE water/steam
			Stage 2  Reagents: Steam/H₂O(g) AND acid/H⁺ (catalyst) ✓		AO3.1	
			+ H <sub>2</sub> O		AO2.6	For steam, ALLOW H <sub>2</sub> O with temperature ≥100°C ALLOW use of H <sub>3</sub> PO <sub>4</sub> /H <sub>2</sub> SO <sub>4</sub> as catalyst DO NOT ALLOW HCI IGNORE pressure

20	(f)	Structure of ester product $\checkmark$ $C_6H_5$ $C_6H_$	2	AO3.1 AO3.2	ALLOW any combination of skeletal OR structural OR displayed formula as long as unambiguous  ALLOW  C <sub>6</sub> H <sub>5</sub> C <sub>6</sub> H <sub></sub>
		Total	22		

Question	Answer	Marks	AO element	Guidance
21	Refer to marking instructions on page 5 of mark scheme for guidance on marking this question.  Level 3 (5–6 marks)  Describes, in detail, reactions of two aliphatic compounds that form a C–C bond  AND mechanisms for the two aliphatic reactions.  There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.  Level 2 (3–4 marks)  Describes a reaction of one aliphatic compound that forms a C–C bond with few omissions/errors.  AND mechanism for one aliphatic reaction.  OR  Describes reactions of two compounds that forms a C–C bond AND attempts a mechanism for one of the reactions  There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.  Level 1 (1–2 marks)	6	AO1.2 ×4 AO2.5 ×2	Indicative scientific points may include:  Reactions of aliphatic compounds and mechanisms  • Haloalkane, RX and CN⁻ → RCN + X⁻ Reagents: NaCN and ethanol Reaction: Nucleophilic substitution Mechanism:  H  R  C  C  R  C  C  C  N  • Aldehyde or ketone and HCN e.g. RCHO + HCN → RCH(OH)CN Reagents: NaCN and H⁺ Reaction: Nucleophilic addition Mechanism:
	Selects suitable reagents for reactions of <b>two</b> compounds that form a C–C bond.  OR  Attempts to describe a reaction and mechanism of <b>one</b> compound that forms a C–C bond, with omissions/errors.  There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.  O marks No response or no response worthy of credit.			R  C  C  C  C  C  C  C  C  C  C  C  C  C
	Total	6		

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