



Mark Scheme (Results)

November 2021

Pearson Edexcel GCE

In Chemistry (9CH0)

Paper 3: General and Practical Principles in
Chemistry

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
 - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
 - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
 - iii) organise information clearly and coherently, using specialist vocabulary when appropriate.

Using the mark scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit. () means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer. ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

Quality of Written Communication

Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

Full marks will be awarded if the candidate has demonstrated the above abilities. Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

Question Number	Answer	Additional Guidance	Mark												
1(a)(i)	<ul style="list-style-type: none"> all numbers for ^{35}Cl correct all numbers for $^{37}\text{Cl}^-$ correct 	<p><u>Example of table</u></p> <table border="1"> <thead> <tr> <th>Particle</th><th>Protons</th><th>Neutrons</th><th>Electrons</th></tr> </thead> <tbody> <tr> <td>^{35}Cl atom</td><td>17</td><td>18</td><td>17</td></tr> <tr> <td>$^{37}\text{Cl}^-$ ion</td><td>17</td><td>20</td><td>18</td></tr> </tbody> </table> <p>If no other mark is awarded, allow (1) for any four numbers correct</p>	Particle	Protons	Neutrons	Electrons	^{35}Cl atom	17	18	17	$^{37}\text{Cl}^-$ ion	17	20	18	(2)
Particle	Protons	Neutrons	Electrons												
^{35}Cl atom	17	18	17												
$^{37}\text{Cl}^-$ ion	17	20	18												

Question Number	Answer	Additional Guidance	Mark
1(a)(ii)	<ul style="list-style-type: none"> lines at 70 and 72 and 74 relative abundances 9:6:1 	<p><u>Example of spectrum</u></p> <p>Allow any abundances in an approximate 9:6:1 ratio e.g. 56:37-38:6 as %, or 75 : 50 : 8</p>	(2)

Question Number	Answer	Additional Guidance	Mark
1(b)	<ul style="list-style-type: none"> KClO_3 	Allow K^+ClO_3^-	(1)

Question Number	Answer	Additional Guidance	Mark
1(c)	<ul style="list-style-type: none"> equation state symbols 	<p>(1) <u>Example of equation</u> $\text{Cl(g)} + \text{e}^- \rightarrow \text{Cl}^-\text{(g)}$ Allow just e for electron</p> <p>(1) Stand alone mark for species on both sides of equation Ignore state symbol for electron</p>	(2)

Question Number	Answer	Additional Guidance	Mark
1(d)(i)	<ul style="list-style-type: none"> identification of oxidising agent 	Either acidified (potassium) manganate(VII) / MnO_4^- and H^+ Or acidified hydrogen peroxide / H_2O_2 and H^+ Allow H^+ shown in equation in (i) or (ii) If the acid is specified it must be sulfuric acid	(1)

Question Number	Answer	Additional Guidance	Mark
1(d)(ii)	<ul style="list-style-type: none"> value of E^\ominus_{cell} 	Either $E^\ominus_{\text{cell}} = (+)0.15 \text{ (V)}$ for acidified (potassium) manganate(VII) Or $E^\ominus_{\text{cell}} = (+)0.41 \text{ (V)}$ for acidified hydrogen peroxide No TE on any other reagent in (i)	(1)

(Total for Question 1 = 9 marks)


Question Number	Answer	Additional Guidance	Mark								
2(a)	A description that makes reference to the following points:	<u>Examples of reagents and observations</u>	(2)								
	• reagent	(1)									
	• observation	(1)									
		<table><tr><th>Reagent</th><th>Observation</th></tr><tr><td>any carbonate / NaHCO_3 / KHCO_3 (added to aqueous acid)</td><td>effervescence / fizzing / bubbles / gas evolved that turns limewater milky</td></tr><tr><td>magnesium (added to aqueous acid)</td><td>effervescence / fizzing / bubbles / gas evolved that gives a pop with a lighted splint</td></tr><tr><td>alcohol and (concentrated) H_2SO_4 / HCl / H^+</td><td>characteristic smell (of an ester)</td></tr></table>	Reagent	Observation	any carbonate / NaHCO_3 / KHCO_3 (added to aqueous acid)	effervescence / fizzing / bubbles / gas evolved that turns limewater milky	magnesium (added to aqueous acid)	effervescence / fizzing / bubbles / gas evolved that gives a pop with a lighted splint	alcohol and (concentrated) H_2SO_4 / HCl / H^+	characteristic smell (of an ester)	
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magnesium (added to aqueous acid)	effervescence / fizzing / bubbles / gas evolved that gives a pop with a lighted splint										
alcohol and (concentrated) H_2SO_4 / HCl / H^+	characteristic smell (of an ester)										
		Allow names or formulae for reagents but if both are given, both must be correct									
		Ignore conditions e.g. heat									
		Do not award PCl_5 / Na									
		If more than one test is given, penalise any incorrect tests									

Question Number	Answer	Additional Guidance	Mark														
2(b)	<p>A description that makes reference to two of the following points:</p> <ul style="list-style-type: none">reagentcorresponding observation	<p><u>Examples of reagents and observations</u></p> <table><tr><th>Reagent</th><th>Observation</th></tr><tr><td>bromine water Allow bromine (in an organic solvent)</td><td>orange / yellow / brown solution goes colourless Allow bromine water is decolourised</td></tr><tr><td>carboxylic acid and (concentrated) H₂SO₄ / HCl / H⁺</td><td>characteristic smell (of an ester)</td></tr><tr><td>acidified potassium manganate(VII) / permanganate</td><td>purple to colourless / decolourised</td></tr><tr><td>alkaline potassium manganate(VII)</td><td>purple to green</td></tr><tr><td>(neutral) potassium manganate(VII)</td><td>purple to brown ppt</td></tr><tr><td>acidified (potassium) dichromate((VI)) (ions)</td><td>orange to green</td></tr></table> <p>Allow names or formulae for reagents but if both are given, both must be correct</p> <p>Ignore conditions e.g. heat</p> <p>Do not award PCl₅ / Na</p> <p>If more than one test is given, penalise any incorrect tests</p>	Reagent	Observation	bromine water Allow bromine (in an organic solvent)	orange / yellow / brown solution goes colourless Allow bromine water is decolourised	carboxylic acid and (concentrated) H ₂ SO ₄ / HCl / H ⁺	characteristic smell (of an ester)	acidified potassium manganate(VII) / permanganate	purple to colourless / decolourised	alkaline potassium manganate(VII)	purple to green	(neutral) potassium manganate(VII)	purple to brown ppt	acidified (potassium) dichromate((VI)) (ions)	orange to green	(2)
Reagent	Observation																
bromine water Allow bromine (in an organic solvent)	orange / yellow / brown solution goes colourless Allow bromine water is decolourised																
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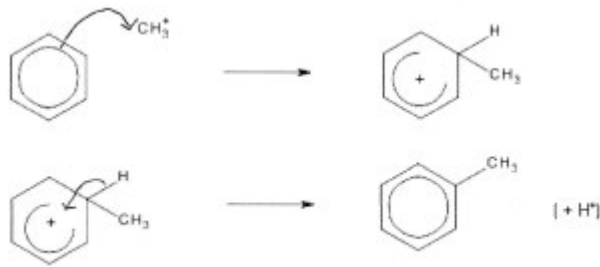
Question Number	Answer	Additional Guidance	Mark
2(c)	<div><div><ul style="list-style-type: none">absorbance for Aabsorbance for Babsorbance for C</div><div><div>(1)</div><div>(1)</div><div>(1)</div></div></div>	<div><div><div><div><div>Absorbance</div><div>Wavenumber range / cm⁻¹</div></div><div><div>Absorbance expected in infrared spectrum of A but not in B or C</div><div>1720 – 1700 Allow 1725 - 1700</div></div><div><div>Absorbance expected in infrared spectrum of B but not in A or C</div><div>1669 – 1645</div></div><div><div>Absorbance expected in infrared spectrum of C but not in A or B</div><div>1740 – 1725 Allow 1740 - 1720</div></div></div></div><div><p>Allow values in reverse order e.g. 1700 – 1720 for A</p><p>If single values are given instead of a range award (2) for 3 correct values within the ranges and (1) for 2 correct values</p><p>Do not award a single value that occurs in two ranges i.e 1720-1725</p><p>If no other mark is awarded, allow (1) for: A 3300-2500 and C 2900-2820 / 2775-2700</p></div></div> <div>(3)</div>	

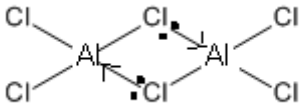
(Total for Question 2 = 7 marks)

Question Number	Answer	Additional Guidance	Mark
3(a)	<ul style="list-style-type: none"> $(1s^2)2s^22p^63s^23p^1$ 	<p>Allow numbers of electrons as subscripts or large numbers</p> <p>Allow p orbitals designated as x, y and z</p> <p>Ignore $1s^2$ repeated</p>	(1)

Question Number	Answer	Additional Guidance	Mark
3(b)(i)	<ul style="list-style-type: none"> dot-and-cross diagram 	<p><u>Example of diagram</u></p>  <p>Allow electrons in overlapping circles</p> <p>Allow all dots / all crosses</p> <p>Ignore inner shell electrons, even if incorrect</p> <p>Ignore lines as bonds e.g.</p> <p style="text-align: center;"> \underline{x} \bullet </p> <p>Do not award diagram with lone pair on Al</p>	(1)

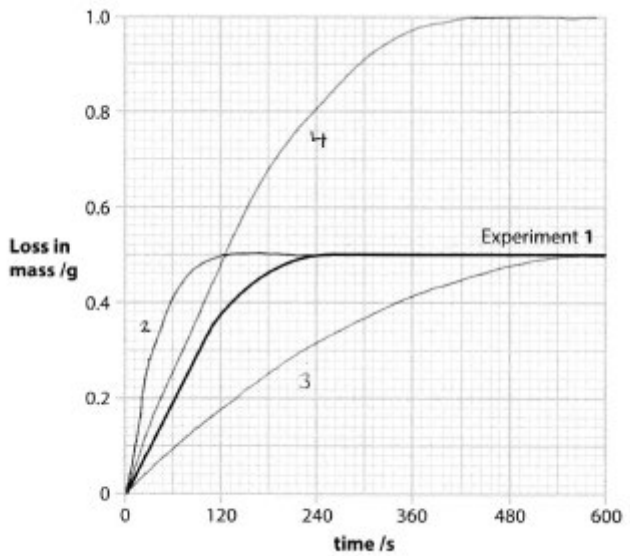
Question Number	Answer	Additional Guidance	Mark
3(b)(ii)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none">• shape – trigonal planar (1)• bond angle - 120° (1)	<p>Mark independently</p> <p>Both words needed Allow triangular for trigonal – but not just tri</p> <p>Allow marks for labelled diagram</p> <p>Note If shape is pyramidal, no mark for M1 but allow (1) for 107° No TE for any other shape</p>	(2)

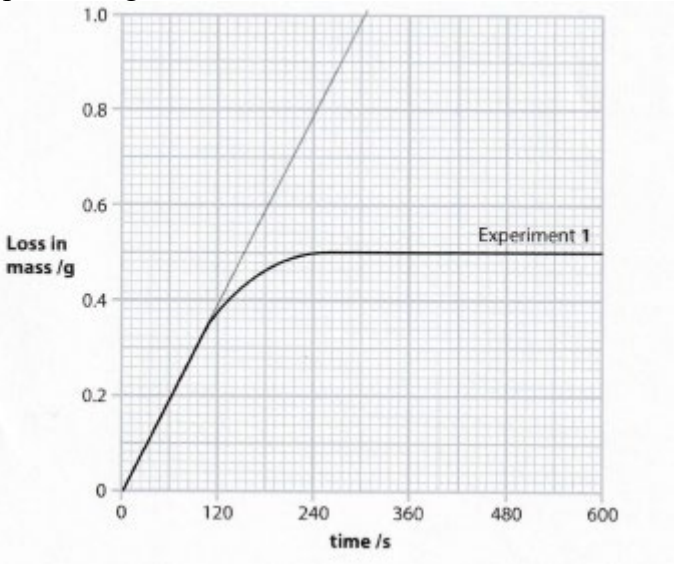
Question Number	Answer	Additional Guidance	Mark
3(b)(iii)	<ul style="list-style-type: none"> equation for the formation of the electrophile (1) curly arrow from on or within the circle to CH_3^+ (1) structure of intermediate including charge with some part of the charge within the horseshoe and horseshoe covering at least 3 carbon atoms and facing the tetrahedral carbon (1) curly arrow from C-H bond to anywhere in the hexagon reforming the delocalised structure (1) 	<p><u>Example of mechanism</u></p> <p>$\text{CH}_3\text{Cl} + \text{AlCl}_3 \rightarrow \text{CH}_3^+ + [\text{AlCl}_4]^-$</p>  <p>Allow $\text{AlCl}_4^- / \delta^+\text{CH}_3\text{-AlCl}_4^{\delta-}$</p> <p>Allow curly arrow from anywhere within the hexagon</p> <p>Allow curly arrow to any part of CH_3^+, including the + charge</p> <p>Do not award curly arrow from outside the hexagon</p> <p>Allow dotted / dashed lines for horseshoe</p> <p>Do not award dotted bonds to H and CH_3 unless clearly part of a 3D structure</p> <p>Ignore any involvement of AlCl_4^- in the final step /HCl</p> <p>Note</p> <p>Correct Kekulé structures score full marks</p>	(4)

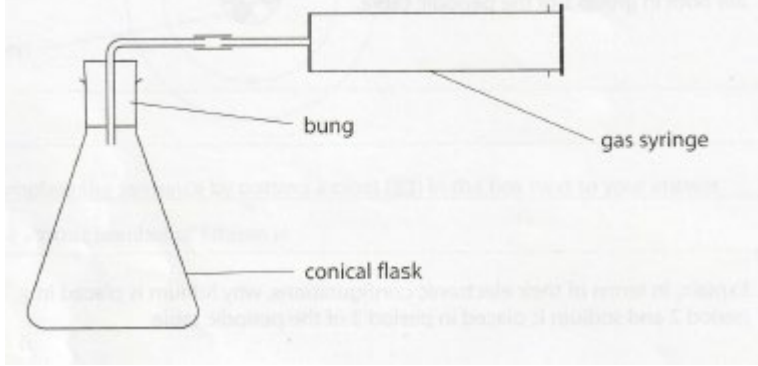
Question Number	Answer	Additional Guidance	Mark
3(c)(i)	<ul style="list-style-type: none"> diagram showing two AlCl_3 molecules joined through two chlorine atoms 	<p>Example of diagram</p>  <p>Allow dot-and-cross diagram</p> <p>Ignore missing arrows / direction of arrows</p> <p>Ignore missing lone pairs</p>	(1)

Question Number	Answer	Additional Guidance	Mark
3(c)(ii)	<ul style="list-style-type: none"> dative (covalent) bonds or coordinate bonds 	<p>Allow this labelled on diagram in (i)</p> <p>Do not award this mark if dative bonds shown as arrows starting from aluminium in (c)(i)</p>	(1)

(Total for Question 3 = 10 marks)

Question Number	Answer	Additional Guidance	Mark
4(a)	<ul style="list-style-type: none"> • curve for Experiment 2 steeper than Experiment 1 and Experiment 4 and finishing at the same mass loss as Experiment 1 (1) • curve for Experiment 3 shallower than Experiment 1 and finishing at the same mass loss (1) • curve for Experiment 4 steeper than Experiment 1 and finishing at mass loss 1.0 g (1) 	<p><u>Examples of curves</u></p>  <p>Allow curve for Experiment 4 steeper than Experiments 1 and 2 and finishing at mass loss 1.0 g</p>	(3)

Question Number	Answer	Additional Guidance	Mark
4(b)	<ul style="list-style-type: none"> • tangent drawn to curve when time = 0 tangent must touch curve for at least first 3 small squares on x axis and extend to at least 120 s (1) • calculation of gradient (1) • units (1) 	<p><u>Example of tangent</u></p>  <p>gradient = $\frac{1.0}{300} = 3.33 \times 10^{-3}$ Allow 3.13×10^{-3} to 3.53×10^{-3}</p> <p>TE on tangent drawn or measurements from line on graph with no tangent</p> <p>Ignore SF including 1SF</p> <p>g s^{-1} or g/s</p> <p>stand alone mark</p>	(3)

Question Number	Answer	Additional Guidance	Mark
4(c)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> a diagram showing (calcium carbonate in a conical) flask attached to a gas syringe / a delivery tube passing into a container of water with an upturned measuring cylinder <p>(1)</p> <ul style="list-style-type: none"> add the hydrochloric acid and (immediately) stopper the flask <p>(1)</p> <ul style="list-style-type: none"> record the volume of gas <p>(1)</p> <ul style="list-style-type: none"> collected at regular time intervals <p>(1)</p>	<p><u>Example of diagram</u></p>  <p>Ignore missing labels Ignore heat / water bath Do not award inclusion of a condenser Do not award test tube or beaker for collecting gas</p> <p>Allow carbonate added to acid and stopper the flask Allow acid in a tube / beaker in the flask and tip the flask for them to mix</p> <p>Allow specified time intervals Allow collected in a specified time</p>	(4)

(Total for Question 4 = 10 marks)

Question Number	Answer	Additional Guidance	Mark
5(a)(i)	<ul style="list-style-type: none"> (The cation in X is) Fe^{2+} / iron(II) / Fe(II) (1) (The anion in X is) SO_4^{2-} / sulfate(VI) (1) 	Allow Fe^{+2} Allow sulfate / SO_4^{-2} Do not award sulfite / sulfate(IV)	(2)

Question Number	Answer	Additional Guidance	Mark
5(a)(ii)	<ul style="list-style-type: none"> species and balancing (1) state symbols (1) 	<u>Examples of equation</u> $\text{Fe}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_2(\text{s})$ or $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_2(\text{s}) + 6\text{H}_2\text{O}(\text{l})$ or $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Fe}(\text{OH})_2(\text{H}_2\text{O})_4(\text{s}) + 2\text{H}_2\text{O}(\text{l})$ Ignore missing square brackets TE on cation that forms an insoluble hydroxide in Test 1 State symbols conditional on correct species or ‘near miss’ / non-ionic equation	(2)

Question Number	Answer	Additional Guidance	Mark
5(a)(iii)	<p>An answer that makes reference to the following point:</p> <ul style="list-style-type: none"> Fe²⁺ is oxidised (to Fe³⁺) by oxygen / air 	<p>Allow iron(III) hydroxide / iron(III) (ions) are formed by reaction with oxygen / air</p> <p>TE on cation in Test 1</p> <p>Allow just 'the precipitate / it is oxidised by oxygen / air'</p>	(1)

Question Number	Answer	Additional Guidance	Mark
5(a)(iv)	<p>An answer that makes reference to the following point:</p> <ul style="list-style-type: none"> to react with / remove any carbonate / sulfite / sulfate(IV) ions or to eliminate the possibility of carbonate / sulfite / sulfate(IV) ions 	<p>Allow to prevent any other ions forming a precipitate with barium ions / Ba²⁺</p>	(1)

Question Number	Answer	Additional Guidance	Mark
5(b)(i)	<ul style="list-style-type: none"> (The cation in Y is) Cu²⁺ / copper(II) (1) (The anion in Y is) Cl⁻ / chloride (1) 	<p>Allow Cu⁺²</p> <p>Ignore water ligands</p> <p>Do not award just copper / Cu</p> <p>Do not award just 'chlorine' / Cl</p>	(2)

Question Number	Answer	Additional Guidance	Mark
5(b)(ii)	<ul style="list-style-type: none"> $[\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$ 	<p>Allow $[\text{Cu}(\text{NH}_3)_4]^{2+}$</p> <p>Allow $[\text{Co}(\text{NH}_3)_6]^{2+}$ if Co^{2+} in (i)</p> <p>Ignore missing square brackets</p>	(1)

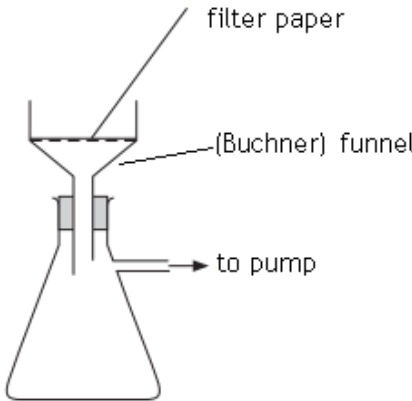
Question Number	Answer	Additional Guidance	Mark
5(b)(iii)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> add dilute (aqueous) ammonia (and stir the mixture) the precipitate dissolves 	<p>(1)</p> <p>(1)</p> <p>If Cl^- / chloride / chlorine ion: Allow add aqueous ammonia / $\text{NH}_3(\text{aq})$ Do not award concentrated ammonia</p> <p>Conditional on use of ammonia Ignore colourless solution</p> <p>If Br^- / bromide / bromine ion identified in Test 4: precipitate is insoluble in dilute ammonia (1) but soluble in concentrated ammonia (1)</p> <p>If I^- / iodide / iodine ion identified in test 4: precipitate is insoluble (1) in dilute and concentrated ammonia (1)</p> <p>Do not award addition of concentrated sulfuric acid</p>	(2)

(Total for Question 5 = 11 marks)

Question Number	Answer	Additional Guidance	Mark
6(a)(i)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> the electron density of the (benzene) ring is greater in phenol (than in benzene) (1) because the lone pair (of electrons) on oxygen and overlaps with the pi cloud / delocalised electrons / delocalised system (1) 	<p>Allow lone pair (of electrons) on oxygen feeds into / donates into / interacts with the delocalised electrons / system</p> <p>Ignore electron pushing effect of OH</p>	(2)

Question Number	Answer	Additional Guidance	Mark
6(a)(ii)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> they both form hydrogen bonds (1) in 4-nitrophenol the hydrogen bonds join molecules in a straight chain / at both ends / at opposite ends (of the molecule so are stronger) or 2-nitrophenol forms intramolecular hydrogen bonds / forces / interactions (so fewer intermolecular hydrogen bonds) (1) 	<p>Allow M1 and M2 shown in diagrams</p> <p>Ignore reference to other specific types of intermolecular forces</p> <p>Allow 4-nitrophenol forms stronger intermolecular hydrogen bonds / forces / interactions</p> <p>Allow in 2-nitrophenol the hydrogen bonds join 2 molecules together / form a dimer (so there are fewer / weaker hydrogen bonds)</p> <p>Allow in 2-nitrophenol the hydrogen bonds are on the same side (of the molecule)</p>	(2)

Question Number	Answer	Additional Guidance	Mark
6(b)	<ul style="list-style-type: none"> reducing agent / reductant 	Ignore tin and concentrated hydrochloric acid Do not award any other named reducing agent	(1)

Question Number	Answer	Additional Guidance	Mark
6(c)(i)	<ul style="list-style-type: none"> side-arm flask with label 'to pump' / drawing of pump (1) (Buchner) funnel with perforations and bung around neck of funnel (1) flat filter paper (over perforations) (1) 	<p><u>Example of diagram</u></p>  <p>filter paper</p> <p>(Buchner) funnel</p> <p>to pump</p> <p>Ignore just 'suction'</p> <p>Allow funnel joined to flask with 'Quickfit' joint / no gap</p> <p>Do not award fluted filter paper / filter paper that extends up the sides of the funnel</p>	(3)

Question Number	Answer	Additional Guidance	Mark
6(c)(ii)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> dissolve crystals in the minimum (amount / volume) (1) of hot water / solvent (1) filter hot and allow to cool (1) filter and wash with a small amount of (cold) solvent (1) dry crystals between filter papers / in a desiccator / in a warm oven (1) 	<p>Allow add / put for dissolve Do not award wash for dissolve</p> <p>Penalise use of incorrect solvent once only</p> <p>Allow M3 if hot is omitted and is mentioned in M2</p> <p>Do not award filter to remove soluble impurities</p> <p>Stand alone mark Allow other suitable methods of drying Do not award reference to crystals mixed with a drying agent</p>	(5)

Question Number	Answer	Additional Guidance	Mark
6(c)(iii)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> melting temperature is lower (1) it melts over a range of temperatures or the melting temperature is not sharp (1) 	<p>Allow a specified range of temperatures</p>	(2)

Question Number	Answer	Additional Guidance	Mark
6(d)(i)	<ul style="list-style-type: none">• $\text{C}_8\text{H}_9\text{NO}_2$	Allow the symbols, with subscripts, in any order e.g. $\text{C}_8\text{H}_9\text{O}_2\text{N}$ Allow large numbers but not superscripts Ignore any other formulae as working	(1)

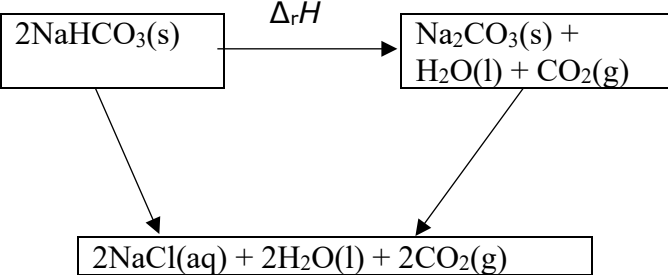
Question Number	Answer	Additional Guidance	Mark
6(d)(ii)	<ul style="list-style-type: none"> calculation of amount of Ce^{4+} (1) calculation of amount of paracetamol in 25.0 cm^3 (1) calculation of amount of paracetamol in 100.0 cm^3 (1) calculation of molar mass of paracetamol and mass of paracetamol in 1 tablet (1) calculation of percentage of paracetamol in 1 tablet and conclusion (1) 	<p><u>Example of calculation</u> amount Ce^{4+} used = $\frac{16.5 \times 0.100}{1000}$ = $0.00165 / 1.65 \times 10^{-3} \text{ (mol)}$</p> <p>amount paracetamol in 25.0 cm^3 = $\frac{0.00165}{2}$ = $0.000825 / 8.25 \times 10^{-4} \text{ (mol)}$ TE on M1</p> <p>amount paracetamol in 100.0 cm^3 = 4×0.000825 = $0.00330 / 3.30 \times 10^{-3} \text{ (mol)}$ TE on M2</p> <p>molar mass of paracetamol, $\text{C}_8\text{H}_9\text{NO}_2$ = $(8 \times 12) + (9 \times 1) + 14 + (2 \times 16) = 151$ and mass of paracetamol in 1 tablet = $0.00330 \times 151 = 0.4983 \text{ g}$ TE on M3 and (d)(i)</p> <p>percentage of paracetamol = $\frac{0.4983}{0.500} \times 100 = 99.66 \text{ (\%)}$ and the tablet was from Brand R TE on M4 provided $M4 < 0.500 \text{ g}$</p> <p>Ignore SF except 1 SF</p>	(5)

(Total for Question 6 = 21 marks)

Question Number	Answer	Additional Guidance		Mark
7(a)(i)	<ul style="list-style-type: none">mass and temperature fall correct			(1)
		Mass of NaHCO ₃ used / g	5.62	
		Temperature fall / °C	(-)6.6	

Question Number	Answer	Additional Guidance	Mark
7(a)(ii)	<ul style="list-style-type: none"> calculation of amount of NaHCO₃ and calculation of amount of hydrochloric acid 	<p>Example of calculation amount NaHCO₃ = $\frac{5.62}{23 + 1 + 12 + (3 \times 16)}$ = 0.0669 (mol) TE on mass of NaHCO₃ in (a)(i) and amount HCl = $\frac{50 \times 2.00}{1000}$ = 0.10 (mol) Ignore SF including 1SF</p>	(2)
	<ul style="list-style-type: none"> 0.0669 mol NaHCO₃ needs 0.0699 mol HCl for reaction so HCl is in excess 	<p>Allow mol ratio = 1 : 1 so HCl is in excess Allow just more moles of HCl used Allow 0.10 > 0.0669 (mol) Allow HCl is in excess by 0.033 (mol)</p>	(1)

Question Number	Answer	Additional Guidance	Mark
7(a)(iii)	<ul style="list-style-type: none"> calculation of heat absorbed (1) calculation of enthalpy change (1) positive sign and units (1) 	<p><u>Example of calculation</u></p> <p>heat absorbed = $50.0 \times 4.18 \times 6.6$ $= 1379.4 \text{ (J)} / 1.3794 \text{ (kJ)}$</p> <p>Ignore sign</p> <p>enthalpy change = $\frac{1379.4}{0.0669}$ $= 20619 \text{ (J mol}^{-1}\text{)}$</p> <p>or $= \frac{1.3794}{0.0669}$ $= 20.619 \text{ (kJ mol}^{-1}\text{)}$</p> <p>TE on heat absorbed and amount NaHCO_3 in (a)(ii)</p> <p>Final answer $+20.6(19) \text{ kJ mol}^{-1}$ or $+20619 \text{ J mol}^{-1}$</p> <p>TE on enthalpy change Allow $+19.7(06) \text{ kJ mol}^{-1}$ from 0.07 mol in (a)(ii) Allow $\text{kJ mol}^{-1} / \text{J mol}^{-1}$</p> <p>Ignore SF except 1 SF Ignore incorrect / missing units in M1 and M2</p> <p>Correct answer with sign and units scores (3)</p>	(3)

Question Number	Answer	Additional Guidance	Mark
7(b)(i)	<ul style="list-style-type: none"> correct species and balancing numbers in lower box (1) both arrows pointing in correct directions (1) 	<p><u>Example of Hess cycle</u></p>  <p>The diagram shows a Hess cycle. At the top left, a box contains the chemical formula $2\text{NaHCO}_3(\text{s})$. A purple arrow labeled $\Delta_r H$ points from this box to a box at the top right containing $\text{Na}_2\text{CO}_3(\text{s}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$. From the bottom of the top-left box, an arrow points down to a box at the bottom containing $2\text{NaCl}(\text{aq}) + 2\text{H}_2\text{O}(\text{l}) + 2\text{CO}_2(\text{g})$. From the bottom of the top-right box, an arrow points down to the same bottom box.</p> <p>Ignore missing state symbols</p> <p>Stand alone mark Ignore labels on arrows and inclusion of HCl</p>	(2)

Question Number	Answer	Additional Guidance	Mark
7(b)(ii)	<ul style="list-style-type: none"> expression for $\Delta_r H$ (1) substitution of values into expression with both values in same units (1) calculation of $\Delta_r H$ and sign and units (1) 	<p><u>Example of calculation</u> $\Delta_r H = 2 \times \Delta H_1 - \Delta H_2$</p> <p>$\Delta_r H = 2 \times 20.619 - (-29.4)$ or $\Delta_r H = 2 \times 20619 - (-29\,400)$ M1 can be scored from values substituted into correct expression in M2 TE on ΔH_1 in (a)(iii) and expression in M1 No TE on incorrect arrows in cycle</p> <p>$\Delta_r H = +70.638 \text{ kJ mol}^{-1}$ or $\Delta_r H = +70638 \text{ J mol}^{-1}$ TE on ΔH_1 in (a)(iii) and expression in M1 provided it is a +ve answer</p> <p>Ignore SF except 1 SF Correct answer with sign and units scores (3)</p>	(3)

Question Number	Answer	Additional Guidance	Mark
7(c)(i)	<ul style="list-style-type: none"> calculation of percentage error 	<u>Example of calculation</u> percentage error $= \frac{(90 - 74)}{90} \times 100 = 17.778 / 17.8 / 18(\%)$ Allow 17.7 recurring Ignore SF except 1SF Do not award 17.7	(1)

Question Number	Answer	Additional Guidance	Mark
7(c)(ii)	<ul style="list-style-type: none"> calculation of percentage uncertainty using measuring cylinder and burette 	<u>Example of calculation</u> percentage uncertainty using measuring cylinder $= \frac{0.5 \times 100}{50} = 1(\%)$ and percentage uncertainty using burette $= \frac{2 \times 0.05}{50} \times 100 = 0.2(\%)$ Ignore SF / ±	(1)

Question Number	Answer	Additional Guidance	Mark
7(c)(iii)	<p>An answer that makes reference to the following point</p> <ul style="list-style-type: none">the difference in the uncertainty in using the burette compared with the measuring cylinder is very much smaller than the % error in the value obtained (so other factors are more significant)	<p>Allow the uncertainty using the burette is not significantly less than using the measuring cylinder</p> <p>Allow uncertainty represents a spread of values whereas the error is the difference of the true value and value obtained</p> <p>Allow just 'hydrochloric acid / HCl is in excess'</p> <p>Ignore heat loss</p>	(1)

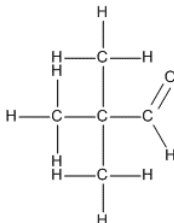
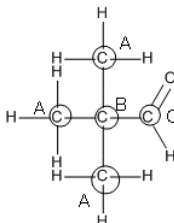
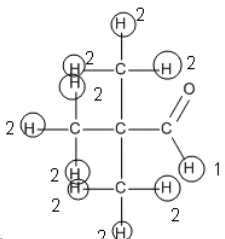
Question Number	Answer	Additional Guidance	Mark
7(c)(iv)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> • measure the temperature of the (hydrochloric) acid every 30 s for 2½ minutes (1) • add the sodium carbonate / solid (at exactly 3 minutes) (1) • (stir and) measure the temperature (of the mixture) every 30 s for another 5 minutes (1) • plot a graph of temperature against time (1) • (join the two sets of points with 2 best fit straight lines and) extrapolate the lines to the time of mixing (1) <p>and determine the maximum temperature change / rise at that time</p>	<p>Allow different times in M1, M2 and M3 or measure the temperature at regular time intervals</p> <p>Allow use of a lid / additional insulation</p> <p>M4 & M5 can be awarded from a suitably labelled sketch graph</p>	(5)

(Total for Question 7 = 19 marks)

Question Number	Answer	Additional Guidance	Mark
8 (a)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> • (add a solution of) iodine and alkali / sodium hydroxide / potassium hydroxide / hydroxide ions (and warm) (1) or • (add a solution of) potassium iodide in sodium chlorate(I) (and warm) • (only) pentan-2-one give a (pale) yellow precipitate / ppt(e) / solid (1) 	<p>Allow names or formulae but if both are given, both must be correct</p> <p>Stand alone mark Allow antiseptic smell</p> <p>Ignore observation for pentan-3-one unless also stated that it gives a yellow precipitate</p>	(2)

Question Number	Answer	Additional Guidance	Mark
8(b)(i)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> curly arrow from lone pair on C of CN^- to C of ketone group (1) curly arrow from $\text{C}=\text{O}$ to, or just beyond, O (1) intermediate (1) curly arrow from lone pair on O^- to H and curly arrow from H-CN bond to anywhere on CN (1) 	<p>Example of mechanism:</p> <p>Allow C_3H_7 and CH_3 for propyl and methyl groups</p> <p>Allow CN bond displayed</p> <p>Ignore correct dipoles, penalise an incorrect dipole once only</p> <p>Do not award M3 if C^+ is shown on intermediate</p> <p>For M4, allow curly arrow from lone pair on O^- to H^+ ion / H_2O molecule</p> <p>Penalise incorrect ketone once only in M3 intermediate</p> <p>Penalise curly arrow from -ve charge instead of lone pair once only</p>	(4)

Question Number	Answer	Additional Guidance	Mark
8(b)(ii)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none">• pentan-2-one / ketone is planar about the carbonyl carbon (1)• so the CN^- / nucleophile attacks (equally) from above and below / either side (of the plane) (1)	<p>Allow bonds about $\text{C}=\text{O}$ are (trigonal) planar or the carbonyl carbon is (trigonal) planar</p> <p>Do not award planar molecule / reference to planar intermediate / ion</p> <p>Do not award multiple directions</p>	(2)

Question Number	Answer	Additional Guidance	Mark
8(c)	<p>An answer that makes reference to the following points:</p> <ul style="list-style-type: none"> displayed formula of aldehyde (1) three different carbon environments indicated (1) two different proton environments indicated (1) no splitting as there are no hydrogens on the adjacent carbon atom(s) (1) 	<p>Example of displayed formula:</p>  <p>Allow CH₃ groups but aldehyde group must be displayed</p> <p>Example of three carbon environments:</p>   <p>Example of two proton environments:</p> <p>Stand alone mark</p>	(4)

Question Number	Acceptable Answers	Additional Guidance	Mark												
8(d)*	<p>This question assesses a student’s ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning.</p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table><tr><td>Number of indicative marking points seen in answer</td><td>Number of marks awarded for indicative marking points</td></tr><tr><td>6</td><td>4</td></tr><tr><td>5–4</td><td>3</td></tr><tr><td>3–2</td><td>2</td></tr><tr><td>1</td><td>1</td></tr><tr><td>0</td><td>0</td></tr></table> <p>The following table shows how the marks should be awarded for structure and lines of reasoning.</p>	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5–4	3	3–2	2	1	1	0	0	<p>Guidance on how the mark scheme should be applied:</p> <p>The mark for indicative content should be added to the mark for lines of reasoning. For example, an answer with five indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p>	(6)
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points														
6	4														
5–4	3														
3–2	2														
1	1														
0	0														

		Number of marks awarded for structure of answer and sustained line of reasoning	In general it would be expected that 5 or 6 indicative points would get 2 reasoning marks, and 3 or 4 indicative points would get 1 mark for reasoning, and 0, 1 or 2 indicative points would score zero marks for reasoning.	
	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2		
	Answer is partially structured with some linkages and lines of reasoning.	1		
	Answer has no linkages between points and is unstructured.	0		
	Comment: Look for the indicative marking points first, then consider the mark for structure of answer and sustained line of reasoning			

	<p>Indicative content</p> <ul style="list-style-type: none"> • IP1 Reagents and conditions for oxidation – acidified potassium dichromate(VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and $\text{H}^+/\text{H}_2\text{SO}_4$ or acidified sodium dichromate(VI) / $\text{Na}_2\text{Cr}_2\text{O}_7$ and $\text{H}^+/\text{H}_2\text{SO}_4$ or Fehling's solution / Benedict's solution or Tollens' reagent / ammoniacal silver nitrate / $\text{Ag}(\text{NH}_3)_2^+$ • IP2 Oxidation of aldehyde – structure of pentanoic acid • IP3 Oxidation of ketone–pentan-3-one / ketone is not (easily) oxidised • IP4 Reagents and conditions for reduction – lithium tetrahydridoaluminate(III) / lithium aluminium hydride and dry ether /ethoxyethane (followed by a dilute acid) or sodium tetrahydridoborate(III) / sodium borohydride and aqueous / methanol solution (followed by a dilute acid) • IP5 Reduction of aldehyde – structure of pentan-1-ol • IP6 Reduction of ketone – structure of pentan-3-ol 	<p>Reagents - Allow names or formulae but if both are given, both must be correct</p> <p>Products - Allow any combination of displayed and structural formulae / skeletal formulae Allow C_4H_9 / C_3H_7 for the alkyl groups</p> <p>Allow acidified dichromate(VI) ions / $\text{Cr}_2\text{O}_7^{2-}$ and $\text{H}^+/\text{H}_2\text{SO}_4$ Allow acidified manganate(VII) ions / MnO_4^- and $\text{H}^+/\text{H}_2\text{SO}_4$ Ignore reference to heat Do not award just Cu^{2+} for Fehling's / Benedict's</p> <p>$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{COOH}$</p> <p>Allow lithal Ignore hydrogen and platinum (catalyst) Ignore reference to heat</p> <p>$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$</p> <p>$\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{CH}_3$</p>	
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(Total for Question 8 = 18 marks)

Question Number	Answer	Additional Guidance	Mark
9(a)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> there are fewer moles / molecules / particles of (gas) on the right (1) so (equilibrium) yield of ammonia increases (1) 	<p>Any reference to equilibrium constant changing scores (0) overall</p> <p>Allow 4 moles / molecules of gas on the left and 2 moles / molecules on right</p> <p>Allow 'equilibrium shifts to the right' M2 is conditional on M1 or the idea of fewer particles on the right / increasing the value of the quotient / Q</p> <p>Allow reverse argument</p>	(2)

Question Number	Answer	Additional Guidance	Mark
9(b)	<ul style="list-style-type: none"> rearrangement of formula (1) substitution of correct values (1) calculation of K_c (1) units (1) 	<p><u>Example of calculation</u></p> $K_c = K_p \times (RT)^{\Delta n}$ $K_c = 3.55 \times 10^{-2} \times (0.0821 \times 500)^2$ $K_c = 59.821$ <p>TE on Δn</p> <p>Stand alone mark $\text{dm}^6 \text{mol}^{-2}$ or $\text{mol}^{-2} \text{dm}^6$</p> <p>Correct value with units and no working scores (4)</p> <p>Ignore SF except 1 SF</p> <p>M1 and M2 can be in reverse order</p>	(4)

Question Number	Answer	Additional Guidance	Mark																				
9(c)	<ul style="list-style-type: none"> calculation of eqm moles (1) expressions for 3 partial pressures (1) substitution of values into K_p expression (1) rearrangement of K_p expression (1) calculation of total pressure and answer to 1 / 2 SF (1) 	<p><u>Example of calculation</u></p> <table border="1"> <tr> <td></td><td>N₂</td><td>H₂</td><td>NH₃</td></tr> <tr> <td>Initial mol</td><td>1.0</td><td>3.0</td><td>-</td></tr> <tr> <td>Eqm mol</td><td>1.0 - 0.15 = 0.85</td><td>3.0 - (3 x 0.15) = 2.55</td><td>0.30</td></tr> <tr> <td>Total mol at eqm</td><td colspan="3">0.85 + 2.55 + 0.30 = 3.7</td></tr> <tr> <td>Partial pressure</td><td>$\frac{0.85 \times P}{3.7}$</td><td>$\frac{2.55 \times P}{3.7}$</td><td>$\frac{0.30 \times P}{3.7}$</td></tr> </table> $K_p = 7.76 \times 10^{-5} = \frac{\left(\frac{0.30 \times P}{3.7}\right)^2}{\left(\frac{0.85 \times P}{3.7}\right)\left(\frac{2.55 \times P}{3.7}\right)^3}$ $7.76 \times 10^{-5} = \frac{0.087419}{P^2}$ $P^2 = 1126.5 \text{ (atm}^2\text{)}$ $P = 33.564$ $= 34 / 30 \text{ (atm)}$ <p>Allow any symbol for total pressure</p> <p>Allow TE throughout</p> <p>Correct answer to 1 or 2 SF with some working scores (5)</p> <p>Correct answer to 1 or 2 SF with no working scores (4)</p>		N ₂	H ₂	NH ₃	Initial mol	1.0	3.0	-	Eqm mol	1.0 - 0.15 = 0.85	3.0 - (3 x 0.15) = 2.55	0.30	Total mol at eqm	0.85 + 2.55 + 0.30 = 3.7			Partial pressure	$\frac{0.85 \times P}{3.7}$	$\frac{2.55 \times P}{3.7}$	$\frac{0.30 \times P}{3.7}$	(5)
	N ₂	H ₂	NH ₃																				
Initial mol	1.0	3.0	-																				
Eqm mol	1.0 - 0.15 = 0.85	3.0 - (3 x 0.15) = 2.55	0.30																				
Total mol at eqm	0.85 + 2.55 + 0.30 = 3.7																						
Partial pressure	$\frac{0.85 \times P}{3.7}$	$\frac{2.55 \times P}{3.7}$	$\frac{0.30 \times P}{3.7}$																				

Question Number	Answer	Additional Guidance	Mark
9(d)	<ul style="list-style-type: none"> substitution of numbers into expression (1) evaluation of $\Delta H/R$ and $1/T_1 - 1/T_2$ (1) rearrangement of expression (1) evaluation of expression (1) 	<p><u>Example of calculation</u></p> $\ln\left(\frac{K_2}{6.76 \times 10^5}\right) = \left(\frac{-92400}{8.31}\right)\left(\frac{1}{298} - \frac{1}{310}\right)$ $\ln\left(\frac{K_2}{6.76 \times 10^5}\right) = -11119.1 \times 1.299 \times 10^{-4}$ $= -1.4444$ $K_2 = 6.76 \times 10^5 \times e^{-1.4444}$ <p>TE on M2</p> $K_2 = 1.59467 \times 10^5 / 159467(\text{atm}^{-2})$ <p>TE on M3</p> <p>Allow answer from earlier correct rounding to 2 or more SF</p> <p>Ignore SF except 1 SF</p> <p>Correct answer with no / some working scores (4)</p>	(4)

(Total for Question 9 = 15 marks)

