

GCSE (9–1)

Chemistry A (Gateway Science)

J248/04: Paper 4 (Higher Tier)

General Certificate of Secondary Education

Mark Scheme for Autumn 2021

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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1. Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

2. Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
✓	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
<u>—</u>	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

3. Subject-specific Marking Instructions

INTRODUCTION

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

The breakdown of Assessment Objectives for GCSE (9-1) in Chemistry:

	Assessment Objective
AO1	Demonstrate knowledge and understanding of scientific ideas and scientific techniques and procedures.
AO1.1	Demonstrate knowledge and understanding of scientific ideas.
AO1.2	Demonstrate knowledge and understanding of scientific techniques and procedures.
AO2	Apply knowledge and understanding of scientific ideas and scientific enquiry, techniques and procedures.
AO2.1	Apply knowledge and understanding of scientific ideas.
AO2.2	Apply knowledge and understanding of scientific enquiry, techniques and procedures.
AO3	Analyse information and ideas to interpret and evaluate, make judgements and draw conclusions and develop and improve experimental procedures.
AO3.1	Analyse information and ideas to interpret and evaluate.
AO3.1a	Analyse information and ideas to interpret.
AO3.1b	Analyse information and ideas to evaluate.
AO3.2	Analyse information and ideas to make judgements and draw conclusions.
AO3.2a	Analyse information and ideas to make judgements.
AO3.2b	Analyse information and ideas to draw conclusions.
AO3.3	Analyse information and ideas to develop and improve experimental procedures.
AO3.3a	Analyse information and ideas to develop experimental procedures.
AO3.3b	Analyse information and ideas to improve experimental procedures.

Question			Answer	Marks	AO element	Guidance
1			D ✓	1	2.1	
2			A ✓	1	1.2	
3			D ✓	1	1.1	
4			C ✓	1	1.1	
5			C ✓	1	2.2	
6			C ✓	1	2.1	
7			B ✓	1	2.2	
8			B ✓	1	1.1	
9			C ✓	1	1.1	
10			D ✓	1	1.1	
11			B ✓	1	2.2	
12			C ✓	1	1.1	
13			D ✓	1	2.1	
14			C ✓	1	1.1	
15			D ✓	1	1.1	

For answers to Section A if an answer box is blank ALLOW correct indication of answer e.g. circled or underlined.

Question			Answer	Marks	AO element	Guidance
16	(a)		Any three from: Mass spectrum Highest m/z value or molecular ion peak is at 46 which is the M_r of ethanol ✓ Peak at m/z = 31 indicates -CH₂OH group ✓ Peak at m/z = 15 indicates -CH₃ group ✓ Infrared spectrum Idea that IR spectrum shows peak in range 3230-3550 which indicates an O-H bond ✓ Idea that IR spectrum shows peak at approx. 1050 which indicates a C-C bond ✓ Idea that IR spectrum shows peak at just below 3000 which indicates a C-H bond ✓ Idea that IR spectrum shows peak at approx. 1100 which indicates a C-O bond ✓	3	3.1b	ALLOW m/z value linked to any other molecular fragment ALLOW correct link between wavenumber and bond <u>from spectrum</u>
	(b)		Any two from: More sensitive / can analyse very small amounts of substances ✓ More accurate ✓ Faster / can carry out analysis all the time ✓	2	1.1	IGNORE more precise
	(c)	(i)	$\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$ Formulae ✓ Balancing ✓	2	1.1 2.1	ALLOW any correct multiple, including fractions DO NOT ALLOW and / & instead of '+' balancing mark is dependent on the correct formulae but ALLOW 1 mark for a balanced equation with a minor error in subscripts / formulae e.g. $\text{C}_2\text{H}_5\text{Oh} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$

Question			Answer	Marks	AO element	Guidance
	(c)	(ii)	Produces soot / produces carbon monoxide / produces less energy ✓	1	1.1	ALLOW produces a toxic or poisonous gas IGNORE produces a harmful gas
	(d)		FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 61 / 60.9 / 60.87 (%) award 2 marks Atom economy = $\frac{28.0}{(28.0 + 18.0)} \times 100$ / $\frac{28.0}{46.0} \times 100$ ✓ = 61(%) / 60.9 (%) / 60.87(%) ✓	2	2.1	ALLOW atom economy formula in words for one mark i.e. atom economy = $\frac{\text{total Mr of desired products}}{\text{total Mr of all products}} \times 100$ ALLOW ECF ALLOW any correct rounding from calculator value, 60.86956522

Question			Answer	Marks	AO element	Guidance
17	(a)	(i)	Idea of looking at each stage of the life of a product to work out the potential environmental impact at each stage ✓	1	1.1	
		(ii)	Any two from: Raw materials needed ✓ Energy used in processing or manufacturing ✓ Water used in processing or manufacturing ✓ Energy needed to <u>use</u> the product ✓ Energy needed to <u>maintain</u> the product ✓ Water or other substances needed to maintain the product ✓ Energy needed to <u>dispose</u> of the product ✓ Space needed to dispose of the product ✓	2	1.1	ALLOW sustainability ALLOW idea of environmental impact of transporting raw materials ALLOW do the materials used decompose or break down ALLOW can the product be recycled IGNORE references to cost IGNORE references to waste products or pollution (stem of question)

Question			Answer	Marks	AO element	Guidance
	(b)	(i)	Vehicle operation ✓	1	3.1a	
		(ii)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 9.5 (tonnes) award 4 marks Mass of CO ₂ produced by petrol car = 80% of 29.8 tonnes = 23.84 (tonnes) ✓ Mass of CO ₂ produced by diesel car = 70% of 20.5 tonnes = 14.35 (tonnes) ✓ Difference = 23.84 – 14.35 = 9.49 (tonnes) ✓ To 2 sig figs = 9.5 (tonnes) ✓	4	3.2b	ALLOW ECF ALLOW ECF ALLOW 1 mark for correct identification of percentages of CO ₂ from vehicle operation for each car (petrol – 80%, diesel – 70%), if no other mark awarded ALLOW answers given to 2 sig figs throughout the question, i.e. Mass of CO ₂ produced by petrol car = 24 (tonnes) Mass of CO ₂ produced by diesel car = 14 (tonnes) Difference = 10 (tonnes)

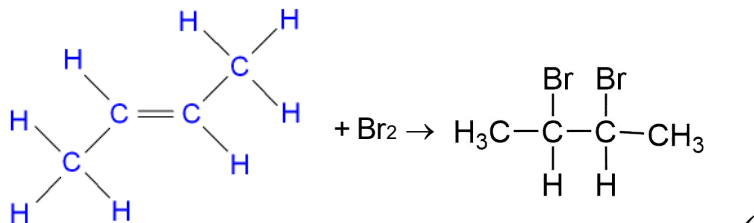
Question			Answer	Marks	AO element	Guidance
18	(a)		Use a gas syringe ✓	1	3.3b	ALLOW use a burette
	(b)	(i)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 0.31 (cm³/s) award 2 marks Gradient = $\frac{\Delta y}{\Delta x} / \frac{44-7}{120} / \frac{37}{120}$ ✓ = 0.31 (cm ³ /s) ✓	2	2.2 3.1a	ALLOW 0.3 / 0.308 / 0.3083 ALLOW ECF from incorrect calculation of volume ÷ time
		(ii)	Line starting at origin but steeper than original line ✓ Levels off at exactly 40 cm ³ ✓	2	2.2	
	(c)		The equation shows two HCl reacting with one CaCO ₃ / ORA / Mole ratio is 2 HCl : 1 CaCO ₃ / ORA / 1 mol of HCl reacts with 0.5 mol of CaCO ₃ / ORA / ✓	1	2.2	ALLOW 0.5 mol of HCl reacts with 0.25 mol of CaCO ₃ / ORA ALLOW idea that only 0.25 mol of CaCO ₃ reacts ALLOW idea that more moles of HCl are needed so the CaCO ₃ is in excess

Question	Answer	Marks	AO element	Guidance
18 (d) *	Please refer to the marking instructions on page 4 of this mark scheme for guidance on how to mark this question. Level 3 (5–6 marks) Analyses the results to correctly identify the effects of changing the temperature AND changing the concentration, with a correct explanation of the results. <i>There is a well-developed line of reasoning which is clear and logically structured. The information presented is relevant and substantiated.</i> Level 2 (3–4 marks) Analyses the results to correctly identify the effects of changing the temperature OR changing the concentration, with a correct explanation of the results. OR Correctly identifies the effects of changing the temperature AND changing the concentration OR Correctly explains the effect of changing the temperature AND changing the concentration. <i>There is a line of reasoning presented with some structure. The information presented is relevant and supported by some evidence.</i> Level 1 (1–2 marks) Analyses the results to correctly identify the effects of changing the temperature OR changing the concentration. OR Correctly explains the effect of changing the temperature OR changing the concentration. <i>There is an attempt at a logical structure with a line of reasoning. The information is in the most part relevant.</i> 0 marks <i>No response or no response worthy of credit.</i>	6	3 x 1.2 3 x 3.2b	AO3.2b • results of experiments 1 & 2 show that as the temperature increases the reaction time decreases, so the rate of reaction increases • results of experiments 2 & 3 show that as the concentration increases the reaction time decreases, so the rate of reaction increases • reaction rate is fastest in experiment 3 due to the combined effects of increased temperature & increased concentration AO1.2 Increasing concentration: • idea of more crowded particles / particles are closer together / more particles per unit volume / more acid particles per cm³ / more acid particles in the same space • idea of more collisions per second / collisions more often / increased collision frequency / more chance of a collision IGNORE references to just ‘more particles’ Increasing the temperature: • idea that acid particles move faster / particles have more energy • idea of more collisions per second / collisions more often / increased collision frequency / more chance of a collision • idea of more successful collisions / collisions between marble chips and acid are more energetic IGNORE references to ‘faster’ collisions

Question			Answer	Marks	AO element	Guidance
19	(a)		Boiling point of fluorine Answer in range -50 to -200 (°C) ✓ Melting point of astatine Answer in range 150 to 320 (°C) ✓	2	3.1a	
	(b)		$\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$ Balancing ✓ Formulae ✓	2	1.1 2.1	ALLOW any correct multiple, including fractions DO NOT ALLOW and / & instead of '+' balancing mark is dependent on the correct formulae but ALLOW 1 mark for a balanced equation with a minor error in subscripts / formulae eg $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
	(c)		(Group 7) atoms gain electrons ✓ Fluorine atoms are smaller (than astatine) / ORA / idea that outer electron shell is closer to the nucleus in fluorine / ORA / fluorine has less electron shells (than astatine) / ORA less shielding in fluorine (than astatine) / ORA ✓ Electrons are more strongly attracted to fluorine atoms / ORA / fluorine atoms gain electrons more easily / ORA / less energy needed to gain outer electron in fluorine / ORA ✓	3	1.1	Assume unqualified answer refers to fluorine IGNORE fluorine atoms gain electrons more quickly / ORA

Question		Answer	Marks	AO element	Guidance
	(d)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 8.14(g) award 4 marks M_r of $\text{FeCl}_3 = 162.3 \checkmark$ Mass of iron chloride = $\frac{162.3}{55.8} \times 2.80 / \frac{324.6}{111.6} \times 2.80 /$ $2.90 \times 2.80 \checkmark$ $= 8.144(\text{g}) \checkmark$ To 3 sig figs = 8.14 (g) $\checkmark$	4	2.1	ALLOW $2\text{FeCl}_3 = 324.6$ ALLOW ECF from incorrect M_r of FeCl_3 ALLOW ECF ALLOW alternative method, e.g. M_r of $\text{FeCl}_3 = 162.3 \checkmark$ Moles of iron = $\frac{2.80}{55.8} = 0.05 /$ Moles of iron chloride = $0.05 \checkmark$ Mass of iron chloride = 0.05×162.3 $= 8.115 \checkmark$ To 3 sig figs = 8.12 (g) $\checkmark$

Question			Answer	Marks	AO element	Guidance
20	(a)		LPG Petrol Diesel Bitumen All 4 correct = 2 marks Any 2 correct = 1 mark	2	2.1	
	(b)	(i)	Any three from: Idea of high demand for smaller molecules ✓ Smaller molecules used for fuels / alkenes used for polymers ✓ Idea of not enough supply of smaller molecules ✓ Idea of surplus of larger molecules ✓ Idea that cracking changes larger molecules to smaller molecules or alkenes ✓	3	1.1	ALLOW named larger and smaller molecules, e.g. bitumen and LPG / petrol ALLOW molecules with double bonds or unsaturated molecules are used for polymers IGNORE idea that the molecules are too long
		(ii)	Idea that any carbon to carbon bond in the molecule can break ✓	1	2.1	
		(iii)	$\text{C}_{30}\text{H}_{62} \rightarrow \text{C}_{20}\text{H}_{42} + 2 \text{C}_5\text{H}_{10}$ ✓	1	2.1	

Question			Answer	Marks	AO element	Guidance
	(c)	(i)	Any one from: Compound B is an alkene / compound A is an alkane ✓ Compound B is unsaturated / compound A is saturated ✓ Compound B contains a double bond / compound A only has single bonds ✓	1	1.2	Assume unqualified answer refers to compound B
		(ii)	 OR $C_4H_8 + Br_2 \rightarrow C_4H_8Br_2$ ✓	1	2.2	ALLOW structure of -CH₃ group shown in product ALLOW butene drawn as a straight chain or product drawn at an angle as butene ALLOW Br₂C₄H₈

Question			Answer	Marks	AO element	Guidance								
21	(a)			3	1.1	ALLOW named nucleotides, i.e. adenine/thymine/cytosine/guanine ALLOW glucose / monosaccharide								
			<table><tr><td>Polymer</td><td>Type of monomer</td></tr><tr><td>DNA</td><td>Nucleotides ✓</td></tr><tr><td>Protein</td><td>Amino acids ✓</td></tr><tr><td>Starch</td><td>Sugars ✓</td></tr></table>				Polymer	Type of monomer	DNA	Nucleotides ✓	Protein	Amino acids ✓	Starch	Sugars ✓
			Polymer				Type of monomer							
			DNA				Nucleotides ✓							
			Protein				Amino acids ✓							
Starch	Sugars ✓													
	(b)		H C H H C6H5 Double bond between carbon atoms ✓ Rest of structure correct ✓	2	2.1									
	(c)		O HO-C □ C-OH O HO-□-OH ✓ ✓	2	1.1	ALLOW 1 mark if a di-carboxylic acid and di-alcohol are seen								

Question			Answer	Marks	AO element	Guidance
	(d)		Any two from: Avoid naked flames or sources of ignition because solvent is highly flammable ✓ Wear gloves because reagents are corrosive ✓ Wear eye protection because reagents are corrosive ✓	2	1.2 2.2	ALLOW (Carry out experiment) in a well-ventilated lab / fume cupboard ✓ ALLOW a description of 2 precautions for 1 mark if no other mark awarded

Question			Answer	Marks	AO element	Guidance
22	(a)	(i)	150 (dm ³) ✓	1	2.1	
		(ii)	300 (dm ³) ✓	1	2.1	ALLOW ECF from (a)(i), ie 2 x (a)(i)
		(iii)	FIRST CHECK THE ANSWER ON ANSWER LINE If answer = 1.6 (g) award 4 marks Moles of I ₂ = $\frac{150}{24}$ / 6.25 ✓ Mass of I ₂ = $\frac{150}{24} \times 253.8$ / 6.25 x 253.8 / 1586.25 g ✓ Mass of I ₂ in kg = 1.586 / 1.58625 / 1.5863 (kg) ✓ To 1 decimal place = 1.6 (kg) ✓	4	2.2	ALLOW ECF from moles of I ₂ ALLOW ECF from mass of I ₂ in grams ALLOW ECF for 1 decimal place mark

[illegible]

Question			Answer	Marks	AO element	Guidance
23	(a)		Correct identification of sodium / Na^+ (from Test 1) ✓ Correct identification of sulfate / SO_4^{2-} (from Test 2) ✓ Correct formula of compound Y - Na_2SO_4 ✓	3	2 x 2.2 1 x 1.2	Award all 3 marks for Na_2SO_4
	(b)	(i)	Idea that more than one cation gives a white precipitate ✓	1	1.2	ALLOW answers which refer to specific cations that give a white precipitate (Al^{3+} , Ca^{2+} , Mg^{2+} , Zn^{2+} , Pb^{2+}) IGNORE incorrect cation charges DO NOT ALLOW incorrect cations
		(ii)	Idea that the chloride ions could have come from the hydrochloric acid ✓	1	3.2b	
		(iii)	Use nitric acid (instead of hydrochloric acid) ✓	1	1.2	ALLOW use a different acid / do not use hydrochloric acid BUT DO NOT ALLOW sulfuric acid

OCR (Oxford Cambridge and RSA Examinations)
The Triangle Building
Shaftesbury Road
Cambridge
CB2 8EA

OCR Customer Contact Centre

Education and Learning

Telephone: 01223 553998

Facsimile: 01223 552627

Email: general.qualifications@ocr.org.uk

www.ocr.org.uk

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